Remember to adapt your procedure and tables for your assigned ratio from table 1 & 2 below. Prepare all entries in your notebook through 5b as directed in the syllabus and your lab manual with the following change:

*In place of Table 2 in your manual, create Table 2 in Excel and set it up to perform all the necessary calculations (you will need to add columns for absorbance & corrected absorbance). Be sure to title the table since in lab you will print it out and hand it in.*

**Calculate the initial concentrations of xylenol orange (H4Q), Al3+ & H+ for your solution prior to lab. The pH of the buffered solution is 2.00. *Fill in the first row of table 2 with the initial concentrations.***

**Table 1.** Assignments section 002 Th am.

|  |  |  |  |
| --- | --- | --- | --- |
| group/bin # | who? | 6.0 × 10-5 M xylenol orange (mL) | 1.4 × 10-4 M Al(NO3)3 (mL) |
| 1 |  | 10.00 | 10.00 |
| 2 |  | 10.00 | 5.00 |
| 3 |  | 5.00 | 10.00 |
| 4 |  | 10.00 | 10.00 |
| 5 |  | 10.00 | 5.00 |
| 6 |  | 5.00 | 10.00 |
| 7 |  | 10.00 | 10.00 |

**Table 2.** Assignments section 003 Th pm.

|  |  |  |  |
| --- | --- | --- | --- |
| group/bin # | who? | 6.0 × 10-5 M xylenol orange (mL) | 1.4 × 10-4 M Al(NO3)3 (mL) |
| 1 |  | 10.00 | 5.00 |
| 2 |  | 5.00 | 10.00 |
| 3 |  | 10.00 | 10.00 |
| 4 |  | 10.00 | 5.00 |
| 5 |  | 5.00 | 10.00 |
| 6 |  | 10.00 | 10.00 |
| 7 |  | 10.00 | 5.00 |